

Chemistry 163B Winter 2013

Introduction to 2nd Law

Chemistry 163B


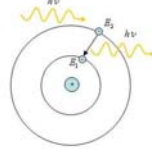

Heuristic Tutorial

Second Law, Statistics and Entropy

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"Lower energy states are more stable"

Systems naturally go to lower energy ????

falling apples
excited state → ground state
chemistry 163B students

http://www.sciencemuseum.org.uk/onlineadults/for_the_saath_a_big_ball_why_dont_we_fall_off_the_bottom_of_k.asp
http://www.thespectroscopynet.com/images/PI_AS_EB_SE.jpg
http://www.moonbattery.com/sleeping_student.jpg

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But ??? First Law of Thermodynamics

The energy lost has to go somewhere ???

Energy lost by system means energy gained by surroundings.

First Law of Thermodynamics

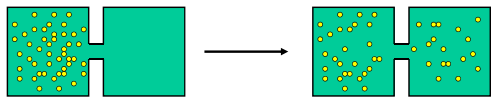
$$\Delta U_{\text{sys}} = -\Delta U_{\text{surr}}$$

$$\Delta U_{\text{universe}} = 0 \quad (\text{also consider } E=mc^2)$$

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The second law of thermodynamics

- Why do things happen?
- What are the limitations on things that do occur spontaneously?
- $\Delta H < 0$ was considered by Berthelot to be the driving principle for spontaneity
- But: some salts cool when dissolving and why do gasses diffuse if no energy difference



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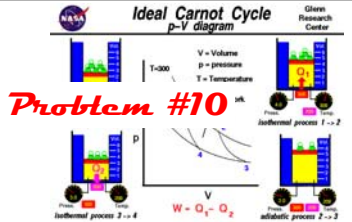
SADI CARNOT

- born 1796, son of Lazare Carnot and antimonarchist
- named Sadi after Persian poet and mystic
- went to Polytechnique Ecole whose faculty included Lagrange, Fourier, Laplace, Berthelot, Ampere, duLong and had as classmates Cauchy, Coriolis, Poisson, Petit, Fresnel
- 1814 went into Corps of Engineers and when monarchy reestablished was sent to boondocks outpost
- 1824 wrote "Reflections on the Motive Power of Heat and Machines Adapted for Developing this Power"

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Carnot cycle (E&R, pp. 80-85 and HW2 prob #10)



Problem #10

1→2 isothermal reversible expansion at T_1 (300K)

2→3 adiabatic reversible expansion $T_1 \rightarrow T_2$ (200K)

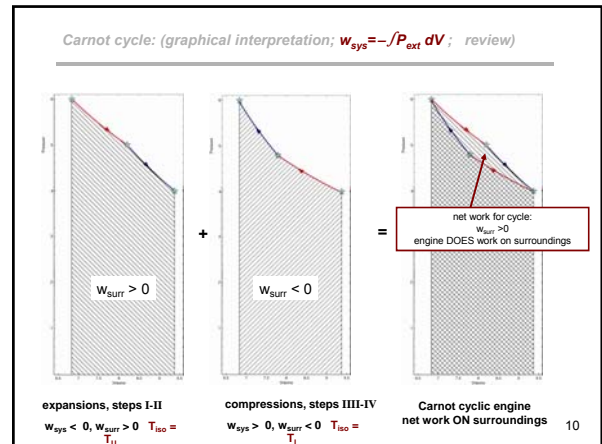
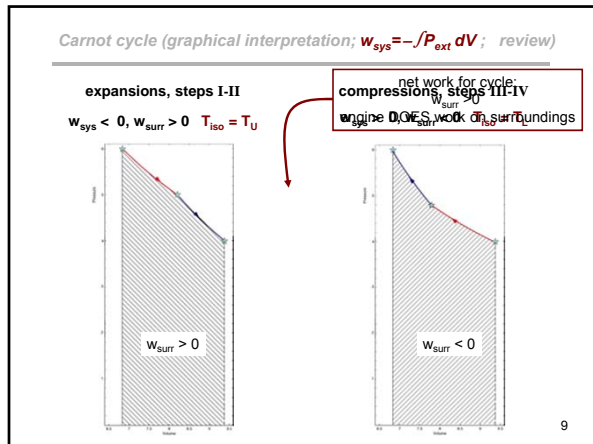
3→4 isothermal compression at T_2 (200K)

4→1 adiabatic compression $T_2 \rightarrow T_1$ (300K)

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from Carnot cycle

for system in complete cycle:
 $\Delta U=0; q > 0; w < 0$ (work DONE on surr) (#10E)

$q > 0$ (q_{in}) at higher T_1 ; $q < 0$ (q_{out}) at lower T_2

efficiency = $-w/q_{1 \rightarrow 2}$
 (how much net work out for heat in 1 \rightarrow 2)

efficiency will depend on T_1 and T_2

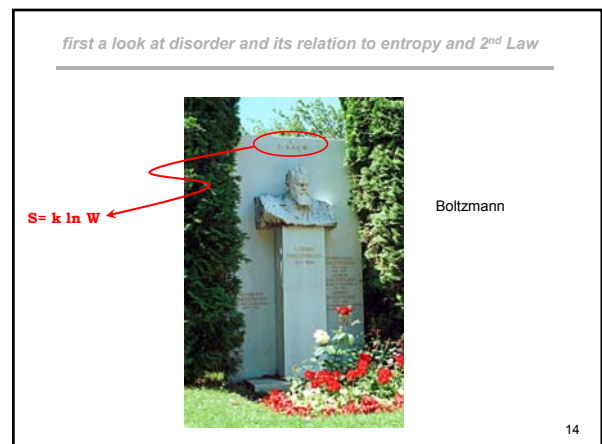
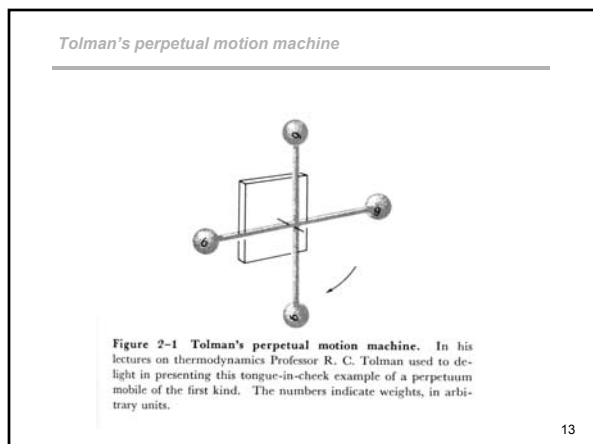
HW prob #22 ϵ is efficiency

$$\epsilon = \frac{T_H - T_C}{T_H} \quad \text{or} \quad \epsilon = \frac{T_U - T_L}{T_U}$$

H=HOT C=COLD or U=UPPER L=LOWER

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- statements of the Second Law of Thermodynamics*
-
1. Macroscopic properties of an isolated system eventually assume constant values (e.g. pressure in two bulbs of gas, becomes constant; two block of metal reach same T) [Andrews. p37]
 2. It is impossible to construct a device that operates in cycles and that converts heat into work without producing some other change in the surroundings. *Kelvin's Statement* [Raff p 157]; *Carnot Cycle*
 3. It is impossible to have a natural process which produces no other effect than absorption of heat from a colder body and discharge of heat to a warmer body. *Clausius's Statement, refrigerator*
 4. In the neighborhood of any prescribed initial state there are states which cannot be reached by any adiabatic process
 ~ Caratheodory's statement [Andrews p. 58]
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disorder and entropy

STATE OF HIGH ENTROPY

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microstates and macrostates

configuration 1 2 3 4	macrostate L R		W no. of microstates

L

R

etc.

$$W = \frac{(n_{\text{total}})!}{(n_L)! (n_R)!}$$

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microstates and macrostates

configuration 1 2 3 4	macrostate L R		W no. of microstates
LLLL	4	0	1
LLLR	3	1	4
LLRL			
LRLL			
RLLL			
LLRR	2	2	6
LRLR			
RLLR			
LRRL			
RLRL			
RRLR			
RRLL			
~	1	3	4
~	0	4	1

L

R

etc.

$$W = \frac{(n_{\text{total}})!}{(n_L)! (n_R)!}$$

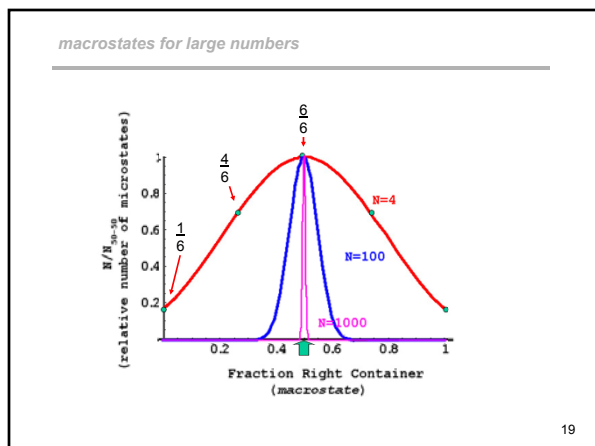
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macrostates and microstates

microstate: one of the **EQUALLY PROBABLE** configurations of molecules (e.g. LLLL vs LRLR)

macrostate: state with specific macroscopic properties
e.g. L=2 R=2

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moral of the story

although each allowed microstate (e.g. LLRR or LLLL) is equally probable

the overwhelming number of microstates correspond to macrostates with almost identical macroscopic properties (e.g. ~ 50-50 RvsL)

W, the number of microstates corresponding to the macrostate, is a measure of the **DISORDER** of the system in that macrostate

a system "meanders" through all available microstates; but you are only likely to observe it in one of the overwhelming number that correspond to the equilibrium macrostate

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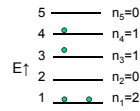
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famous quote

“disorder happens”

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how energy changes affect disorder



n_i = number of molecules in energy state ϵ_i

$$\sum_i n_i = n_{\text{total}}$$

$W = \frac{n_{\text{total}}!}{n_1!n_2!n_3!\dots}$ number of ways of arranging with n_1, n_2, \dots

NOTE : W depends only on the n_i 's, i.e. the distribution of molecules among quantum states

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how energy changes affect disorder

$$U = E_{\text{total}} = \sum_i n_i \epsilon_i$$

$$dE = \sum_i n_i d\epsilon_i + \sum_i \epsilon_i dn_i$$

change in energy due to change in energy levels, e.g. 3D quantum p.i.b. change in energy levels as box changes size

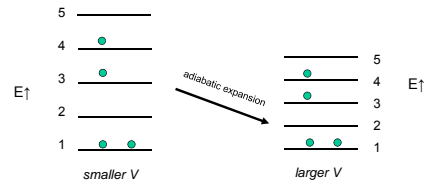
change in energy due to change redistribution of molecules among energy levels, e.g. put in more total energy to fixed size 3D quantum p.i.b.

heat

work

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reversible adiabatic expansion ($\Delta U_{\text{sys}} < 0$; $q_{\text{rev}} = 0$)

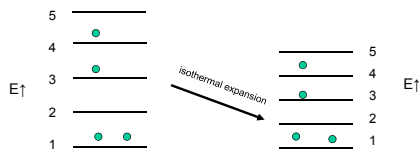


quantum: expansion, bigger box, energy levels more closely spaced

- Total energy of system decreases ($\Delta U_{\text{sys}} < 0$ for adiabatic expansion)
- NO CHANGE IN LEVEL POPULATIONS if expansion done slowly, reversibly
- $q_{\text{rev}} = 0$; NO CHANGE IN DISORDER (W)

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reversible isothermal expansion ideal gas ($\Delta U_{\text{sys}} = 0$; $q_{\text{rev}} > 0$; $w = -q$)

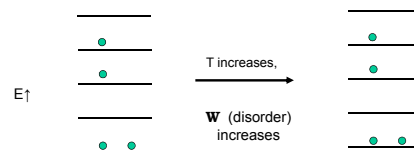


to maintain $\Delta U = 0$ need to put in heat

- Levels get closer due to $\Delta V > 0$; $w < 0$
- To maintain $\Delta U = 0$, $q > 0$ and the level populations have to change and thus W changes
- $q_{\text{rev}} > 0$; INCREASE IN W , INCREASE IN DISORDER

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What happens to W as thermal energy raised?



T increases,

W (disorder) increases

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What happens to W as thermal energy goes to zero (T → 0K) ?

E↑

W decreases, in fact W → 1
only 'one way to distribute at 0°K

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Boltzmann and Entropy

$S = k \ln W$

S is entropy
 k=Boltzmann's constant= $1.3807 \times 10^{-23} \text{ J K}^{-1}$
 disorder increases ⇔ entropy increases

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take home messages

- Disorder, **W**, did not change during an adiabatic reversible expansion ($q_{\text{rev}} = 0$)
- Disorder, **W**, increased in isothermal reversible expansion ($q_{\text{rev}} > 0$)
- Disorder, **W**, increased with T increase ($q > 0$)
- Disorder, **W**, decreased with T decrease ($q < 0$)
- As $T \rightarrow 0$, **W** → 1

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take home messages

A sample problem†:

COLD

$T_A = .63 T_f$

HOT

$T_B = 1.37 T_f$

W (microstates)

$\approx 10^{202}, 10^{354}$

Two bodies 500 molecules each
 one at $T_A = .63 \times T_f$,
 the other at $T_B = 1.37 \times T_f$

bring into thermal contact

STAY COLD-HOT ??

$.63 T_f$
 $1.37 T_f$

$\approx 10^{556} = 10^{202} \times 10^{354}$

or

EQUILIBRIUM ??

$T_E = T_f$

$\approx 10^{594}$

take home message continued†

The equilibrium macrostate is $\frac{10^{594}}{10^{202}} = 10^{392}$ time more likely than the hot-cold state, even though every (microstate)_{hot-cold} has the same likelihood as a (microstate)_{equilibrium}.

No more than one time in 10^{38} a measurement will find the blocks in a half-hot and half-cold configuration.

If you had observed the microstate of the system 10^6 times a second constantly (without a msec of rest!) from the beginning of the universe until your midterm Friday (10^{10} years) **the odds against ever seeing a (microstate)_{hot-cold} are $1:10^{15}$!!!**

† adapted from Nash, *ChemThermo*, Addison Wesley, pp 175-176

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take home message continued†

..... a progressive increase in "disorder" necessarily accompanies an approach to equilibrium characterized by the assumption of macrostates with ever-increasing values of W. And **what may at first appear to be a purposeful "drive" towards states of maximal disorder**, can now be seen to arise from **the operation of blind chance in an assembly** where all microstates remain equally probable, but where the overwhelming proportion of microstates is associated with the maximally disordered (nearly identical) macrostates corresponding to equilibrium macroscopic properties.

† adapted from Nash, *ChemThermo*, Addison Wesley, p 26.

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MUCH MORE

**much more molecules, probability,
statistical mechanics**

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*End Of Introductory Lecture on
Second Law and Disorder*

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Persian Poet 13th Century
Saadi

بی آدم اصحاب یک پیکرند
که در آفرینش یک کونند
هم عضو می بردند آرزوی کارند
که هر عضو را ندانند قرار

literal translation (Farzaneh):
*Of one Essence is the human race
thus has Creation put the base
One Limb impacted is sufficient
For all Others to feel the Mace*

—Saadi (1184–1283)

interpretative translation on UN building:
*Human beings are members of a whole,
In creation of one essence and soul,
If one member is afflicted with pain,
Other members uneasy will remain.
If you have no sympathy for human pain,
The name of human you cannot retain.*



**French Thermodynamicist
Namesake 19th century**
Sadi


and 


$\varepsilon = 1 - \frac{T_L}{T_U}$

heuristic

heu-ris-tic [hyoo-ris-tik or, often, yoo-] **adjective**

1. serving to indicate or point out; stimulating interest as a means of furthering **investigation**.





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