
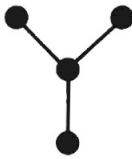
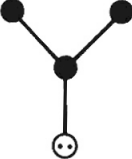
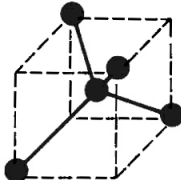
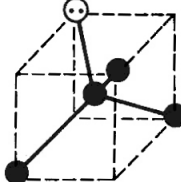
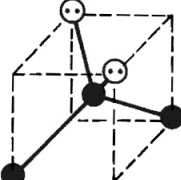
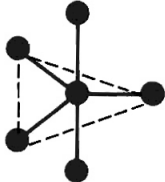
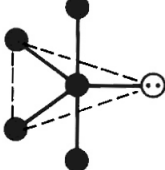
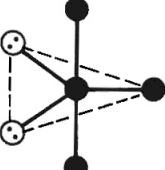
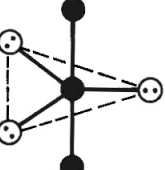
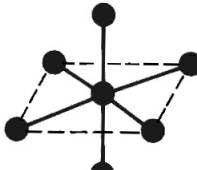
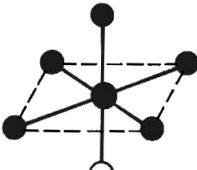
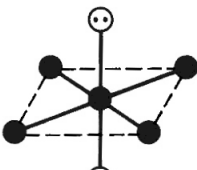


SN	Number of lone pairs	Molecular shape	Example
2	0	 linear	BeH <sub>2</sub> , CO <sub>2</sub>
3	0	 trigonal planar	SO <sub>3</sub> , BF <sub>3</sub>
3	1	 angular	SO <sub>2</sub> , O <sub>3</sub>
4	0	 tetrahedral	CH <sub>4</sub> , CF <sub>4</sub> , SO <sub>4</sub> <sup>2-</sup>
4	1	 trigonal pyramidal	NH <sub>3</sub> , PF <sub>3</sub> , AsCl <sub>3</sub>
4	2	 angular	H <sub>2</sub> O, H <sub>2</sub> S, SF <sub>2</sub>
5	0	 trigonal bipyramidal	PF <sub>5</sub> , PCl <sub>5</sub> , AsF <sub>5</sub>

SN	Number of lone pairs	Molecular shape	Example
5	1	 sawhorse	SF <sub>4</sub>
5	2	 T-shaped	ClF <sub>3</sub>
5	3	 linear	XeF <sub>2</sub> , I <sub>3</sub> <sup>-</sup> , IF <sub>2</sub> <sup>-</sup>
6	0	 octahedral	SF <sub>6</sub> , PF <sub>6</sub> <sup>-</sup> , SiF <sub>6</sub> <sup>2-</sup>
6	1	 square pyramidal	IF <sub>5</sub> , BrF <sub>5</sub>
6	2	 square planar	XeF <sub>4</sub> , IF <sub>4</sub> <sup>-</sup>

from **Chemical Principles 3<sup>rd</sup> Ed**, by Dickerson, Gray, Haight, figure 11-3 (Benjamin, 1979)